

Download Calculate Mole In Compound

Here's a Calculator to Help You. In the case of water, multiply the atomic weight of hydrogen by two, and the atomic weight of oxygen by one, then add the products. Numerically, this would be $(2)(1.008) + (1)(16.00) = 18.016$. This is the molar mass of the compound; it has units of grams per mole. I am trying to calculate the mass Compound A and consequently the volume (as made in my solution containing compound A) that I have to add to the solution containing compound B. However, I must keep the final concentration of the solution at 20mg/mL. Chemists define a mole of a compound as Avogadro's number of molecules of that compound. You can use this information to calculate the number of moles in a sample of a compound with a known weight or mass. Conversely, if you know the number of moles of the compound you have, you can calculate the weight or mass of the sample. Best Answer: A "mole" is just a number. It is a number of something, in this case of molecules (6.02×10^{23}). This is the same for every compound, so a mole of oxygen has the same number of molecules as a mole of carbon dioxide.